

#1 - moles of NaOH?

$$n = M \times V$$

0.1 M buret (L)

#2 - conc. of HCl?

$$M = \frac{n}{V}$$

← (#1) ← (0.02 L)

OR

$$N_A V_A = N_B V_B$$

20 mL 0.1 M buret

(N = M)

$$N_A = \frac{N_B V_B}{V_A}$$

#3 - % error? 0.08 M ↙ average #2

$$\% \text{ error} = \left| \frac{\text{accepted} - \text{experimental}}{\text{accepted}} \right| \times 100\%$$