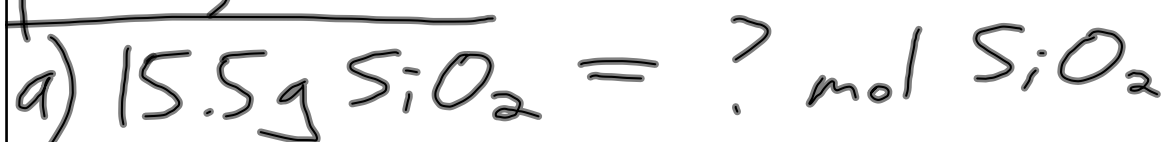


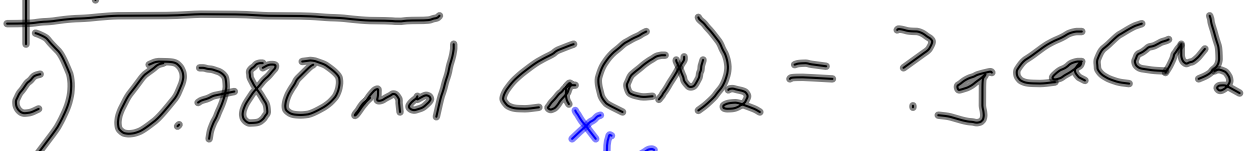
p. 198, # 55



$$\frac{15.5 \cancel{\text{g SiO}_2}}{60.084 \cancel{\text{g SiO}_2}} \times 1 \text{ mol SiO}_2$$

$$= 0.257 \text{ mol SiO}_2$$

p. 198 # 56



$$\frac{0.780 \text{ mol Ca(CN)}_2 \times 92.114 \text{ g Ca(CN)}_2}{1 \text{ mol Ca(CN)}_2}$$

$$= 71.849 \text{ g Ca(CN)}_2$$

$$= 71.8 \text{ g Ca(CN)}_2$$

p.198, #55

d) 5.96 g KOH = ? mol KOH

$$\frac{5.96 \cancel{\text{g KOH}}}{56.104 \cancel{\text{g KOH}}} \times 1 \text{ mol KOH}$$

$$= 0.106 \text{ mol KOH}$$

p.198, #56 72.1498 J/mol

a. 1.50 mol C₅H₁₂ = ? g C₅H₁₂

$$\frac{1.50 \text{ mol C}_5\text{H}_{12} \times 72.1498 \text{ g C}_5\text{H}_{12}}{1 \text{ mol C}_5\text{H}_{12}}$$

$$= 108.225 \text{ g C}_5\text{H}_{12}$$