#1) Strontium to chlorine in strontium chloride?

Sr^{2+} Cl^{-} → Cl_2

$SrCl_2$

$$\frac{Sr}{Cl_2} = \frac{87.62\ g/mol}{(2 \times 35.453\ g/mol)} = 1.24$$

1.24 g Sr : 1 g Cl in $SrCl_2$

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6. sodium bromide

$Na^+ Br^-$

$$\%Na = \frac{Na}{NaBr} \times 100\% = \frac{22.990\ g/mol}{102.894\ g/mol} \times 100\%$$

= 22.34% Na

$$\%Br = \frac{Br}{NaBr} \times 100\%$$ OR

100% - 22.34% = 77.66% Br

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7 males
13 females

males to females?

$\frac{7 \text{ males}}{13 \text{ females}} = 0.54 \text{ males} : 1 \text{ female}$

$\frac{13 \text{ females}}{7 \text{ males}} = 1.86 \text{ females} : 1 \text{ male}$

$\frac{7 \text{ males}}{20 \text{ students}} \times 100\% = 35\% \text{ male}$

$\rightarrow 65\%$ female

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Empirical formulas

- simplified chemical formula
- "least common denominator"

$\frac{4}{6} \Rightarrow \frac{2}{3}$

e.g. $C_6H_{12}O_6$ glucose

$\rightarrow CH_2O$ = empirical formula

- used to identify unknown substances

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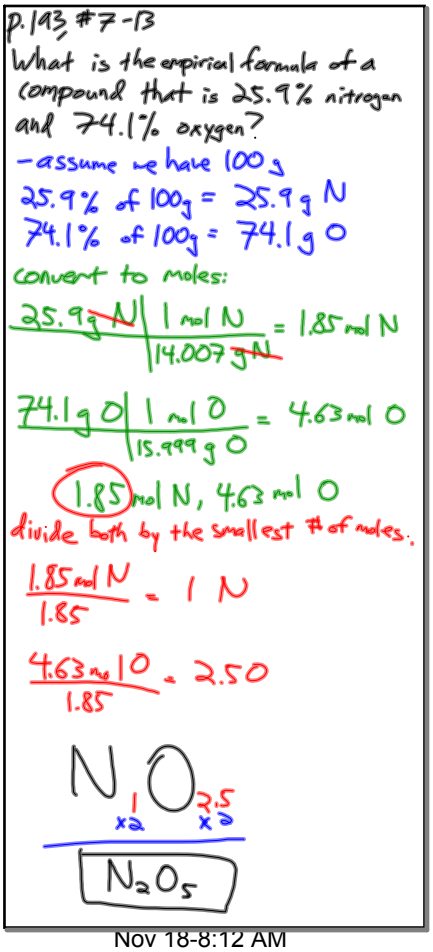

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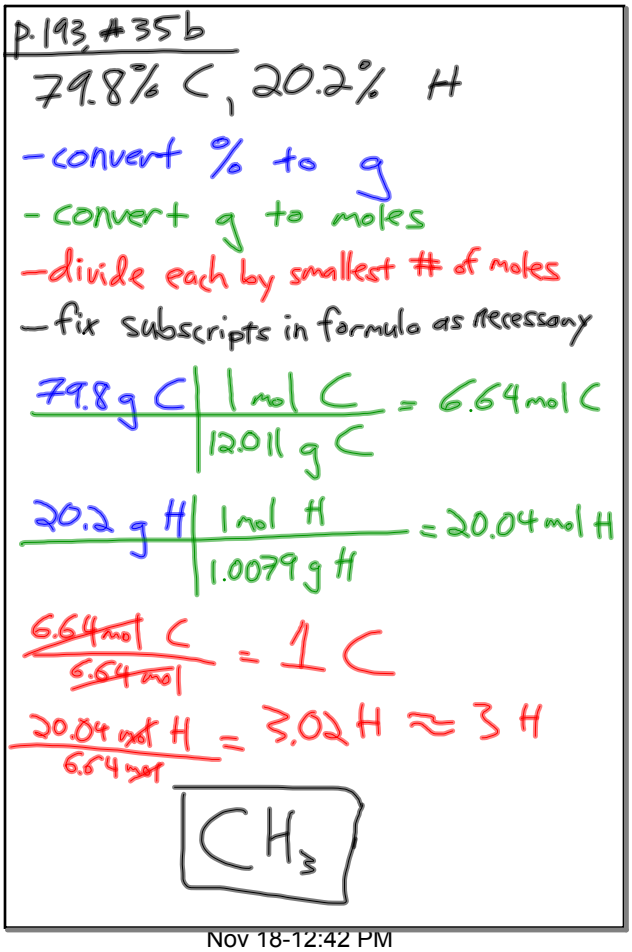

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