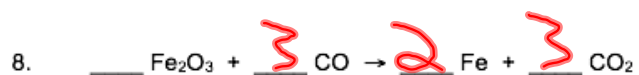
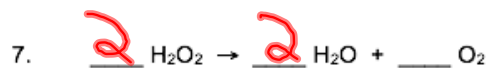
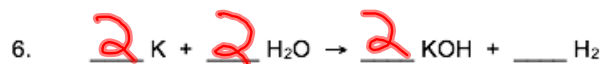
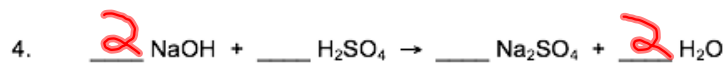
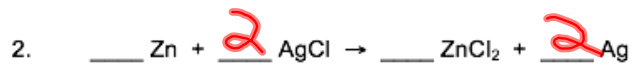


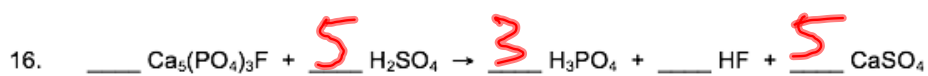
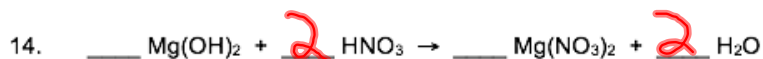
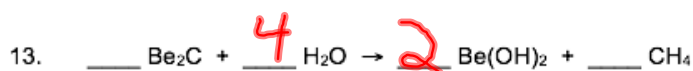
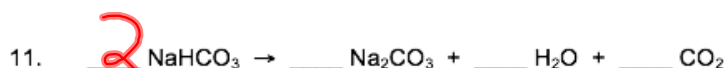
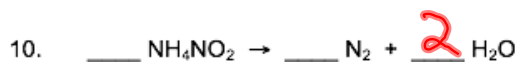
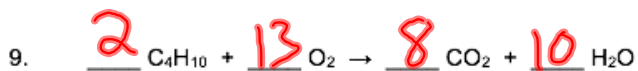
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Date: _____ Per: _____

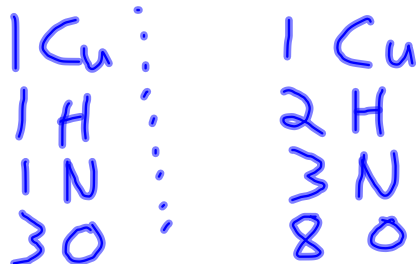
Chemistry
Balancing Equations WS 1

I. Balance the following equations:





Balance the polyatomic ions (PO_4 , SO_4) to make this problem much easier than it looks!



I've set this one up for you, but I'll leave it to you to figure out! Here are some hints:

- because hydrogen is only in water (H_2O) on the right, you should only try even coefficients for HNO_3 on the left
- your final number of nitrogen atoms on either side will be a multiple of 3 (3, 6, 9, etc.)
- balance copper last - trust me!