

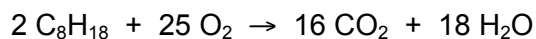
Name: _____

Date: _____

Chemistry
Stoichiometry WS 2

I. Complete the following stoichiometric calculations, balancing equations where necessary.

1. Consider the combustion of octane (C_8H_{18}):

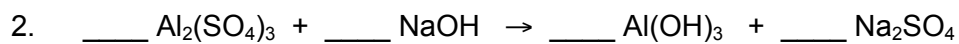


a. How many grams of CO_2 are produced when 191.6 g of octane are burned?

b. How many grams of oxygen gas are required to burn 47.03 g of octane?

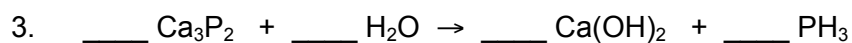
c. How many grams of H_2O are produced when 91.2 g oxygen gas are consumed?

d. How many liters of CO_2 are produced at STP when the reaction yields 5.05 g H_2O ?



a. How many grams of NaOH are needed to completely react with 2.33 g $\text{Al}_2(\text{SO}_4)_3$?

b. If 87.3 g of $\text{Al}(\text{OH})_3$ are formed, how many grams of Na_2SO_4 will be produced?



a. How many grams of water are needed to react with 33.9 g of Ca_3P_2 ?

b. How many grams of PH_3 are produced when the above reaction takes place?

c. How many grams of H_2O will be needed to produce 715 g $\text{Ca}(\text{OH})_2$?