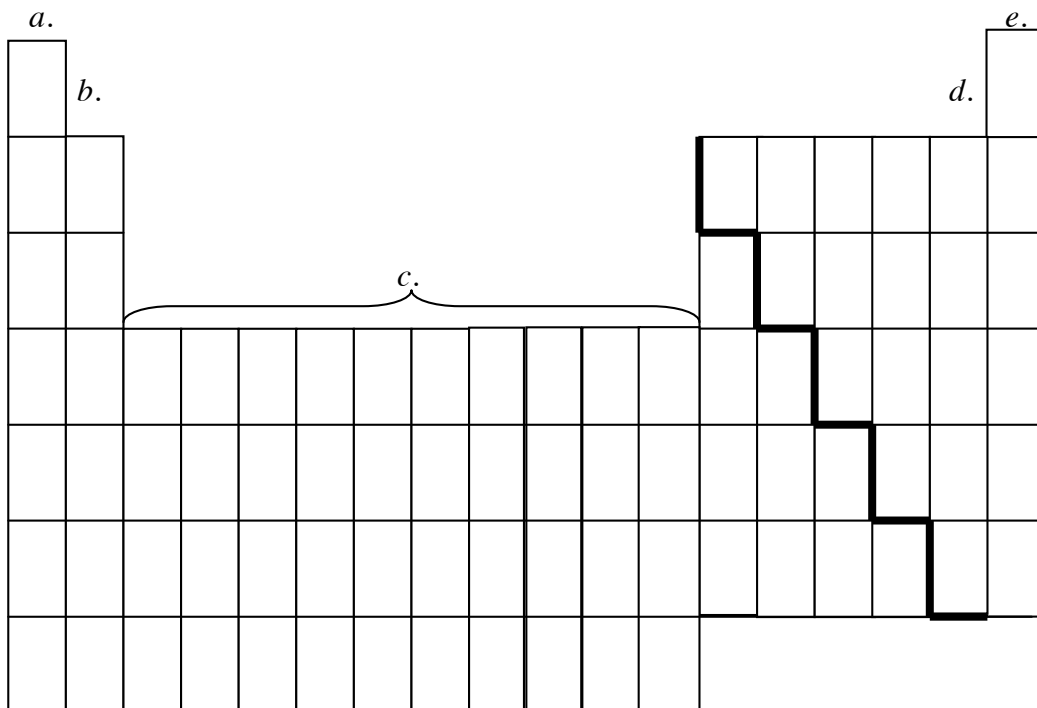


Name: _____

Date: _____ Per: _____

Chemistry Periodicity WS

I. Identify the following groups of the periodic table and give their number of valence electrons:



Group	Group Name	# of valence electrons	ion formed
a			
b			
c			
d			
e			

II. Define the following:

1. periodicity _____

2. electronegativity _____

3. ion _____

4. ionization energy _____

III. *Answer the following questions about periodic trends:*

5. What tends to happen to atomic radius as you go from left to right on the periodic table?

6. What tends to happen to electronegativity as you go from top to bottom on the periodic table?

7. Will boron or aluminum have a larger radius? Why?

8. Will neon or sodium have greater ionization energy? Why?

IV. *Answer the following questions about the element aluminum:*

9. What are three characteristic properties of metals?

10. Aluminum is a metal, but it doesn't exhibit all the normal properties of metals. What property is missing from aluminum? _____

11. Aluminum is next to a dark, jagged line on the periodic table. What is this line called and what does it mean? _____

V. *Answer the following questions about the element magnesium:*

12. What group or family is magnesium found in? _____

13. What is another element from this group? _____

14. How many electrons does magnesium have in its highest energy level? _____

15. What other elements have the same number of outer electrons as magnesium?

16. Knowing their ionization energies, is sodium or magnesium more likely to form an ion? Why?
