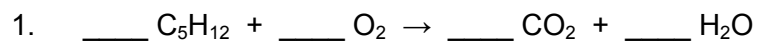


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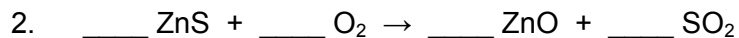
Chemistry
Limiting Reagents WS

I. Complete the following limiting reagent calculations, balancing equations where necessary:



a. 244 g C₅H₁₂ are burned in 160. g O₂. How many grams of H₂O will be formed?

b. In the above reaction, how many grams of carbon dioxide will also be formed?



a. When 75.3 g zinc sulfide are reacted with 217.2 g oxygen gas, how much zinc oxide is formed?

b. In the above reaction, how many Liters of sulfur dioxide will also be formed?

- II. *Complete the following limiting reagent calculations, writing out and balancing the chemical equations:*
3. In a classic neutralization reaction, sodium hydroxide reacts with hydrochloric acid to form sodium chloride (table salt) and water.

a. How much salt would be formed by combining 66.2 g sodium hydroxide with 77.1 g hydrochloric acid?

b. In the above reaction, how many grams of water will also be formed?

4. Iron (III) oxide reacts with carbon monoxide to form iron and carbon dioxide.

a. In this oxidation-reduction reaction, 818.5 g iron (III) oxide are reduced by 134 L carbon monoxide at standard temperature and pressure. How much iron will be formed in this reaction?

b. In the above reaction, how many grams of carbon dioxide will also be formed?